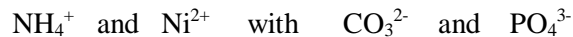


"TRY YOURSELF" QUESTIONS FROM STUDY SECTION 2.5

Try Yourself 2.5

Write formulas for all the ionic compounds that can be formed by combining the following ions:



Try Yourself 2.6

Which compound formula and name in the list is NOT correct?

1. CaSO_4 , calcium sulfate
2. NaNO_3 , sodium nitrate
3. MgI_2 , magnesium iodide
4. NH_4PO_4 , ammonium phosphate
5. $\text{Ca}(\text{ClO})_2$, calcium hypochlorite

Try Yourself 2.7

Sodium oxalate has the formula $\text{Na}_2\text{C}_2\text{O}_4$. Based on this information, what is the formula for iron(III) oxalate?

1. FeC_2O_4
2. $\text{Fe}(\text{C}_2\text{O}_4)_2$
3. $\text{Fe}(\text{C}_2\text{O}_4)_3$
4. $\text{Fe}_2(\text{C}_2\text{O}_4)_3$
5. $\text{Fe}_3(\text{C}_2\text{O}_4)_2$

Try Yourself 2.8

Which of the following are correct formulas for compounds? For those that are not, give the correct formula.

- a) AlCl
- b) NaF_2
- c) Ga_2O_3
- d) MgS
- e) CaO
- f) SrCl_2
- g) $\text{Fe}_2\text{O}_3 / \text{FeO}$
- h) K_2O

Try Yourself 2.9

Which one of the following two ionic compounds melting point will be higher? Explain your answer.

