"TRY YOURSELF" PROBLEMS FROM STUDY SECTION 3.4; 3.5 AND 3.6

Try Yourself 3.6

Predict the products of each precipitation reaction and then balance the complete equation

Lead(II)nitrate(aq) + Potassium bromide(aq) \rightarrow ? Calcium nitrate(aq) + Potassium fluoride(aq) \rightarrow ? Sodium carbonate(aq) + Copper(II) chloride(aq) \rightarrow ? Nickel(II) chloride(aq) + Potassium hydroxide(aq) \rightarrow ?

Try Yourself 3.7

Write a balanced net ionic equation for the following reaction: Aqueous solutions of calcium chloride and sodium phosphate are mixed.