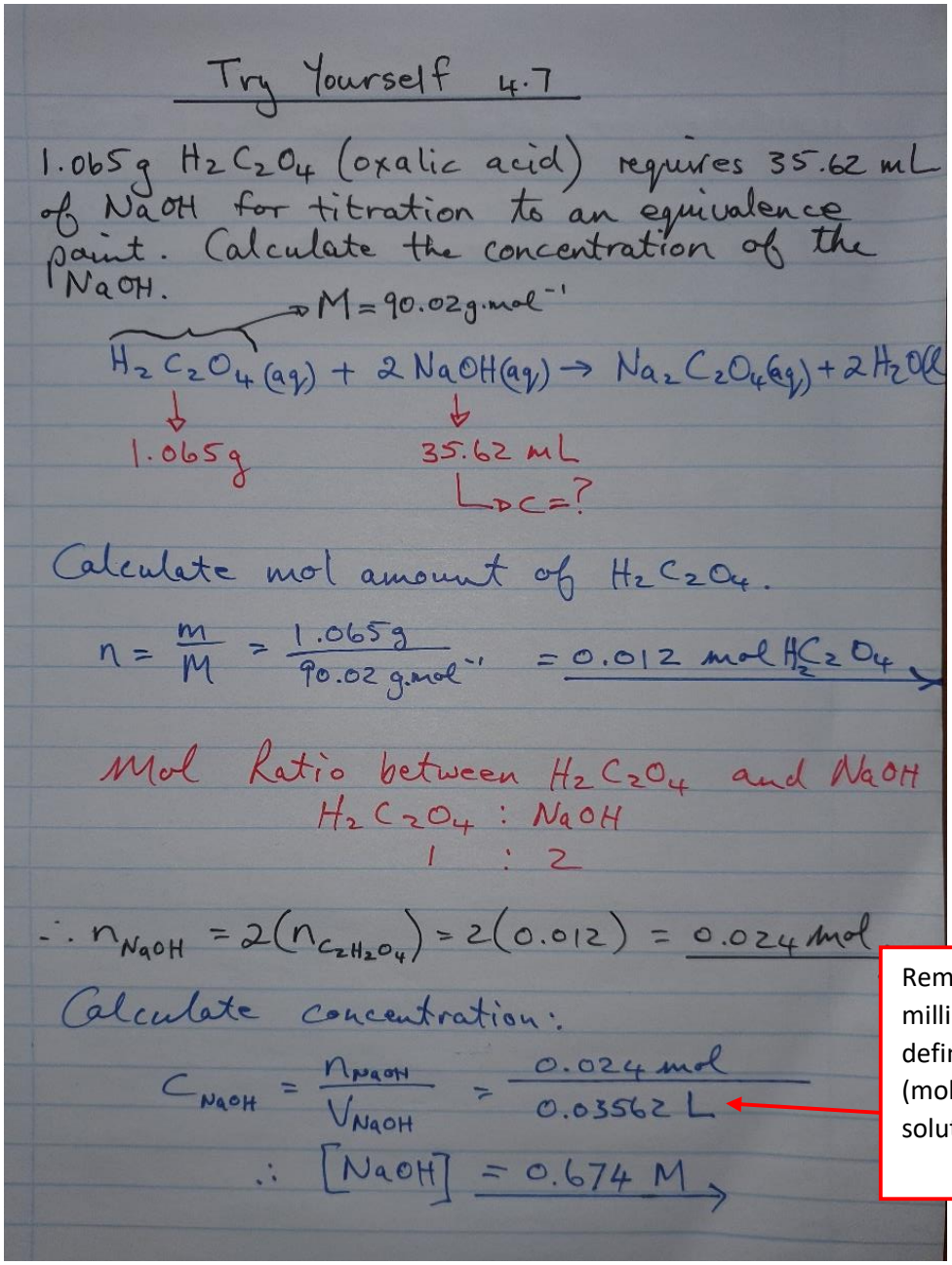
ANSWER TO "TRY YOURSELF" PROBLEM FROM STUDY SECTION 4.7

1.065 g of $H_2C_2O_4$ (oxalic acid) requires 35.62 mL of NaOH for titration to an equivalence point. Calculate the concentration of the NaOH.

$$H_2C_2O_4(aq) + 2\ NaOH(aq) \rightarrow Na_2C_2O_4(aq) + 2\ H_2O(liq.)$$



Remember to convert the volume in millilitre (mL) to dm^3 (L) because per definition the concentration (molarity) is equal to mol per dm^3 solution. So, 35.62 mL = 0.03562 dm^3