"TRY YOURSELF" PROBLEM FROM STUDY SECTION 4.7

Try yourself 4.7

1.065~g of $H_2C_2O_4$ (oxalic acid) requires 35.62~mL of NaOH for titration to an equivalence point. Calculate the concentration of the NaOH.

$$H_2C_2O_4(aq) + 2 \text{ NaOH}(aq) \rightarrow Na_2C_2O_4(aq) + 2 H_2O(liq.)$$