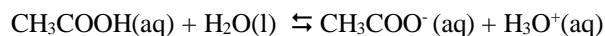
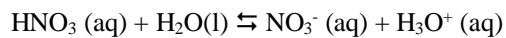


"TRY YOURSELF" PROBLEMS FROM STUDY SECTION 8.2

Try Yourself 8.2 a



Try Yourself 8.2 b

What is the conjugate base of the HCO_3^- ion?

- (1) H_3CO_3^+ (2) HCO_3^- (3) CO_3^{2-}

Which of the following is NOT an acid-base conjugate pair?

- (1) HClO and Cl^- (2) HNO_2 and NO_2^-
(3) HF and F^- (4) H_2CO_3 and HCO_3^-

Try Yourself 8.2 c

Complete this table by identifying the correct conjugate acid or conjugate base.

Acid	Conjugate base	Base	Conjugate acid
HCOOH		CN^-	
H_2S			HSO_4^-
	ClO^-		H_2SO_3

Try Yourself 8.2 d

Write a balanced equation for the reaction that occurs when H_3PO_4 , phosphoric acid, donates a proton to water to form the dihydrogen phosphate ion. Is the dihydrogen phosphate ion an acid, a base or amphiprotic?