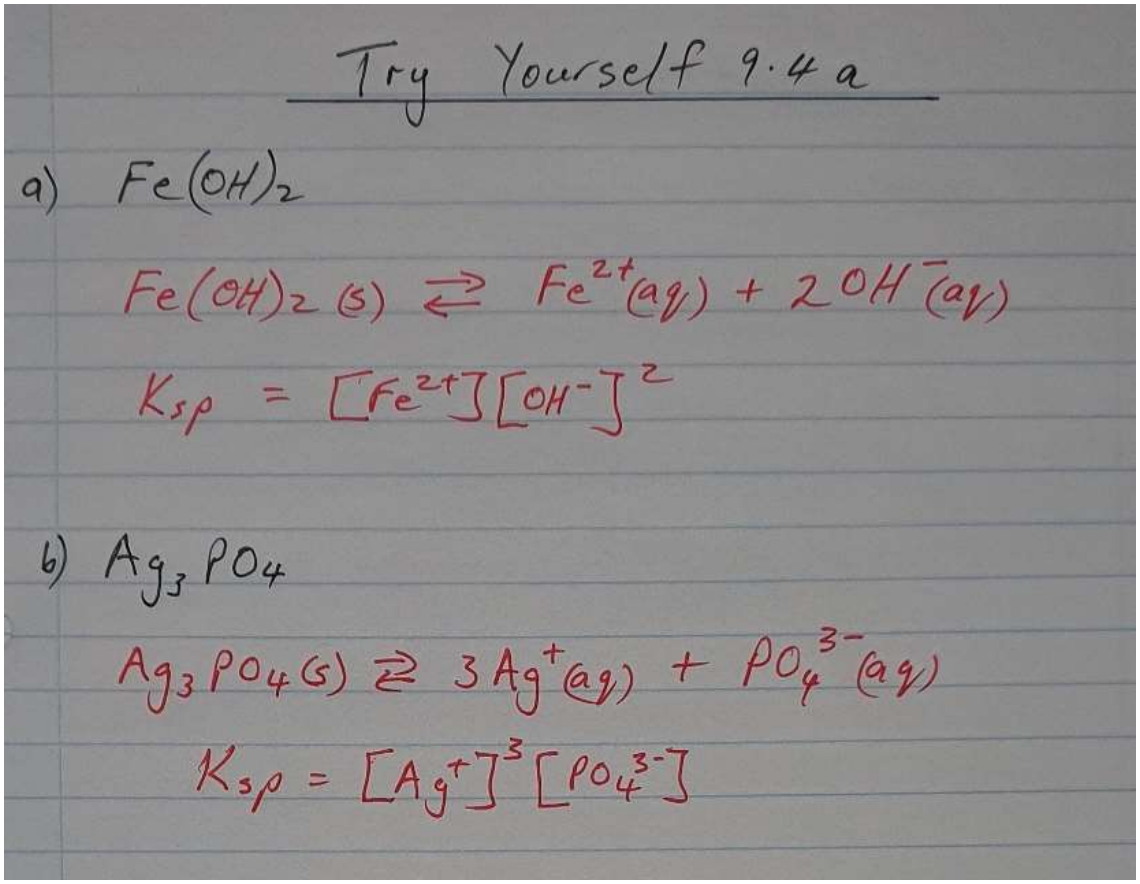
ANSWER TO "TRY YOURSELF" PROBLEM FROM STUDY SECTION 9.4

Try Yourself 9.4a

Write balanced reaction equations and the K_{sp} expression for each of the following slightly soluble salts:

a. $Fe(OH)_2$

b. Ag_3PO_4



Try Yourself 9.4b

Calculate the solubility of MgF_2 in moles per liter and in grams per liter. K_{sp} of $MgF_2 = 5.2 \times 10^{-11}$ and $M_{MgF2} = 62.3\ g.mol^{-1}$