

**"TRY YOURSELF" PROBLEM FROM STUDY SECTION 9.4**

**Try Yourself 9.4a**

Write balanced reaction equations and the  $K_{sp}$  expression for each of the following slightly soluble salts:

- a.  $\text{Fe}(\text{OH})_2$
- b.  $\text{Ag}_3\text{PO}_4$

**Try Yourself 9.4b**

Calculate the solubility of  $\text{MgF}_2$  in moles per liter and in grams per liter.  $K_{sp}$  of  $\text{MgF}_2 = 5.2 \times 10^{-11}$  and

$M_{\text{MgF}_2} = 62.3 \text{ g}\cdot\text{mol}^{-1}$