

"TRY YOURSELF" PROBLEM FROM STUDY SECTION 9.5

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Solid PbI_2 ($K_{\text{sp}} = 9.8 \times 10^{-9}$) is placed in a beaker of water. After a period of time the lead(II) concentration is measured and found to be $1.1 \times 10^{-3} \text{ M}$. Is the solution saturated? If not, will more PbI_2 dissolve?